



S3 Block 1 Chemistry in Society- Pupil Progress Check

	Learning Outcomes	✓😊	?😐	✗😞
1	I can state the definition of a fossil fuel			
2	I can state 4 examples of fossil fuels			
3	I can describe how oil and coal are formed			
4	I can explain what the key words finite, renewable, non-renewable, fuel, combustion and hydrocarbon mean.			
5	I can state that crude oil is a mixture of hydrocarbons			
6	I know that the boiling point of a hydrocarbon depends on its size			
7	I can describe the link between length of hydrocarbon and boiling point			
8	I can describe the process of fractional distillation as the separation of different groups of hydrocarbons according to their boiling points			
9	I can label a diagram of a fractionating tower to show the boiling points of the different fractions and their uses.			
10	I can state that different fractions have different properties - e.g, flammability, volatility, and viscosity			
11	I know that many of the crude oil fractions are burnt to release energy.			
12	I can state that a reaction that gives out energy is called exothermic.			
13	I can identify the products of combustion of a hydrocarbon from experiment.			
14	I can state the tests for carbon dioxide and water			
15	I can represent the combustion of hydrocarbons as word equations			
16	I explain why the demand for the combustion of hydrocarbons has increased			
17	I can describe how the increase in combustion of hydrocarbons has contributed to the			

	greenhouse effect.			
18	I can label photosynthesis, respiration and combustion on a carbon cycle diagram			
19	I can explain why the levels of carbon dioxide in the atmosphere are increasing.			
20	I can state that the 'carbon footprint' is the sum of all CO ₂ emissions of an individual or group over a year.			
21	I can carry out experiments with metals to deduce a reactivity series.			
22	I can list key properties of metals.			
23	I can state examples of metal uses and explain why they are suitable for their use.			
24	I recognise that iron is widely used despite its corrosion.			
25	I can suggest why iron is so widely used.			
26	I can set up a simple cell to generate voltage			
27	I can carry out experiments to find out changing the metal affects the voltage in a simple cell using copper as a reference			
28	I can identify other factors that could affect the voltage of a simple cell			
29	I can analyse experimental data to generate an electrochemical series.			
30	I can state that current flows from the metal higher in the ECS to the lower metal.			
31	I can use the ECS to make predictions about the direction of current flow and which metal combinations will generate the highest voltage.			
32	I can state that a battery is a number of cells combined to create a higher voltage.			
33	I can explain why batteries go flat.			
34	I can state that rechargeable batteries use an electric current to reverse the chemical reaction and regenerate the reactants.			
35	I can compare the advantages and disadvantages of lead-acid and lithium ion batteries.			